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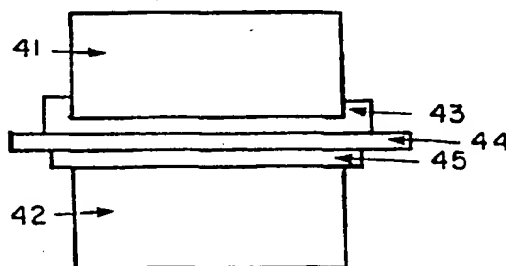


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(54) Title: SOLID ELECTROLYTE AND BATTERIES

*apparently no gel
polymer electrolyte*



(57) Abstract

Novel composite solid electrolytes (CSE) (44) based on a combination which comprises micron-size non-conductive oxide particles, an alkali metal salt, an electronically insulating elastomer and up to a certain percentage of an aprotic solvent. There are provided primary and secondary cells (15) which contain the composite electrolyte. There is also provided an electrochemical cell having an alkali metal transference number lower than 0.5.

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SOLID ELECTROLYTE AND BATTERIES

FIELD OF THE INVENTION:

The present invention relates to improvements in electrochemical cells, and especially in such cells which have an anode which consists of or contains an alkali metal, and preferably lithium.

The invention relates more specifically to composite solid electrolytes (CSE) for use in such electrochemical cells, which CSEs have an alkali metal transference number of at least 0.5. The invention further relates to electrochemical cells, and preferably rechargeable cells which contain such solid electrolytes.

The preferred cathode materials for these cells are FeS_2 , CoS_2 , NiS_2 and their mixtures. The combination of these cathodes with the CSEs gives cells with an electrochemical overcharge protection mechanism.

BACKGROUND OF THE INVENTION:

Solid alkali metal halides of the formula MX , where M is an alkali metal and X is a halide have a moderate ionic conductivity, and a low electronic conductivity.

Conductivity takes place through M^+X vacancy mechanism. An order of magnitude of conductivity can be gained (10^{-5} mho cm^{-1} at ambient temperature) by the mixing of LiI and Al_2O_3 powders and pressing into pellet form. Possible explanations are: (a) interfacial Li^+ conduction; (b) the increase of Li^+ vacancies concentration due to the intimate presence of Al^{3+} at the LiI interface, the Li^+

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conduction being mainly at the $\text{LiI}/\text{Al}_2\text{O}_3$ interface. On this basis an all solid $\text{Li}/\text{LiI}-\text{Al}_2\text{O}_3/\text{PbI}_2$ primary battery was developed.

However the $\text{LiI}-\text{Al}_2\text{O}_3$ pellets are very brittle and have little mechanical strength and are sensitive to thermal shock.

Attempts to use Al_2O_3 with polymers to improve the mechanical strength, in conjunction with lithium salts resulted in a too low Li transference number as too low salt concentrations or too large Al_2O_3 particles or unsuitable salts were used.

Known polymer electrolytes (PEs) have a number of drawbacks, such as a lithium transference number of only about 0.3 to 0.5 which leads to high concentration polarization and interface problems. Frequently salt anions such as BF_4^- , AsF_6^- and ClO_4^- were used which are not compatible with lithium. The above results in a high interfacial Li/PE resistance and lithium passivation. Furthermore the mechanical strength is not adequate, especially at elevated temperatures. The drawbacks set out above deterred the development of high-power high energy alkali metal (M) batteries using polymer electrolytes. Reactivity of the prior art anions with lithium or other alkali metals results in a thick passivating layer, having a high resistance. Furthermore polymer electrolytes tend to soften and start to flow at temperatures in the 100°C range.

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Recently, Nagasubramanian, Peled, Attia and Halpert, ECS Meeting, Toronto, Canada, October 1992, fabricated CSE films made of 12-17% 0.05μ Al_2O_3 particles, 40-55% LiI and 35-53% Polyethylen Oxide (PEO). At $100^\circ C$ the transference number of lithium cation in these films is 0.8-1.0, depending on the film composition.

The state of the art Li/PE batteries which use cathodes such as V_2O_5 , V_3O_{13} , $Li_{0.5}MnO_2$ do not have an internal electrochemical overcharge protection mechanism. Thus it is not practical to stack these cells in a series combination in a bipolar configuration.

SUMMARY OF THE INVENTION:

The present invention relates to improvements in electrochemical cells, and especially in rechargeable cells which have an alkali metal anode. The invention relates to composite solid electrolyte (CSE) which have an alkali metal transference number preferably greater than about 0.5. The invention further relates to electrochemical cells and other devices in which such CSEs can be used. More particularly one aspect of the present invention relates to a composite solid electrolyte (CSE), for use in electrochemical cells, which CSE has an alkali metal

transference number larger than about 0.5, which CSE comprises up to about 40 volume-% electronically non-conductive oxide particles, of an average size of less than about 5μ , which are compatible with alkali metals, up to about 70 volume-% of an alkali metal salt MX or its solid complex with a polymers where M is an alkali metal and X an anion, which is believed to form a thin coating of said oxide particles or is in contact with these, which salt MX is not reducible by the alkali metal in such cell, up to about 80 volume-% of an electronically insulating elastomer, or an hydrocarbon polymer, which elastomer forms a complex with the alkali metal salt, and optionally up to about 20 volume-% of an aprotic organic solvent or low MW polymer, excluding CSE films of polyethylene oxide (PEO), LiI and Al_2O_3 only.

Preferably the CSE according to the invention contains oxide particles in the 0.05 to 0.5μ size range and their quantity is preferably from 1 to 20 volume-% with the quantity of alkali metal salt being from 10 to 40 volume-%. Preferred is an elastic CSE which contains from about 30 to 70 volume-% elastomer and/or hydrocarbon polymer.

Preferably the particles used are particles of oxides of Al, Si, Mg, Ca, Ba, Sr or mixtures of any of these, which may contain some Li, Na, K.

Preferably the alkali metal of the MX salt is Li, Na or K and the anion X is I^- , Br^- , Cl^- , S^- , N_3^- , $CF_3SO_3^-$, $(CF_3SO_2)_2N^-$ or a mixture of any of these.

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A wide variety of elastomers can be used. Preferably the elastomer is a polyalkylene oxide, possibly in combination with one or more of polymethylmethacrylate (PMMA), polymethylacrylate (PMA), polybutylacrylate (PBA), polybutylmethacrylate (PBMA), polyacrylonitrile (PAN), a polyalkylene glycol dialkyl ester or ether, where the ratio of the polyalkylene oxide to the other polymers varies from 50: 1 up to 1:5. Polymers of choice are a mixture of polyethylene oxides of a MW in the 10^5 to 10^7 range where the n to MX ratio preferably is from 1 to 1.5 up to 1 to 20 (n in $P(EO)_n$). The solid CSEs preferably contain an electrolyte such as of up to 20 weight-% diglyme, triglyme or tetraglyme or of a crown ether or polyalkylene glycol dialkyl ether (or ester) of MW 500 to 50,000 or an aprotic organic solvent. Solvents of choice are propylene carbonate (PC), ethylene carbonate (EC), -butyrolactone, DMSO, Diethylene carbonate (DEC), THF, 2MTHF, Dioxolane, DME, or a mixture of these. The CSEs may comprise a MX salt doped by up to 10% atomic ratio of a polyvalent cation selected from Al, Ca, Mg, Sr and Ba cations.

According to further embodiment of the invention there are provided electrochemical cells where t_{Li^+} may be lower than 0.5 which cells are immune to overcharge i.e. they have an internal electrochemical overcharge protection mechanism. These cells consist of PEO based CSE with EO to MX ratio of 6-40 where MX must be MI or a mixture of MI + MX with MCF_3SO_3 and/or $M(CF_3SO_2)_2N$ and a FeS_2 , NiS_2 or CoS_2 cathode or a mixture of these.

The invention further relates to secondary electrochemical cells comprising a solid composite electrolyte as above described.

Cells according to the invention contain an anode made from an alkali metal (M) or alkali metal alloy or a M_xC anode, where M_xC is an alkali metal - carbon or graphite insertion compound where $x < 0.2$; a composite solid electrolyte sheet, as described, compatible with the anode, and a cathode, which is easily charged and discharged against the relevant alkali metal and current collectors for the anode and the cathode, with the preferred M in M_xC being Li.

Preferred are batteries where the alkali metal is lithium, the CSE consists of lithium iodide admixed with up to about 50 mole percent MI or LiX. Amongst suitable cathode materials are: TiS_2 , Li_xMnO_2 ($0.5 < x < 1$), V_2O_5 , CoS_x , NiS_x ($0.8 < x < 4.5$), V_3O_{16} , FeS or FeS_2 .

The cathode material may be mixed with carbon and CSE components to form a composite cathode.

The cell can be assembled with a cathode in the discharge state. For example $2Fe + Li_2S$, $2Ni + Li_2S$, $2Co + Li_2S$, $Fe + Li_2S$, $Li_7V_3O_{16}$ etc. These are the final discharge product of the cell. Thus fine powders must be well mixed and cast with the composite polymer, instead of using charged state of the cathode.

These procedures are preferred in the case of using carbon or graphite anode, which on charge receives lithium.

For batteries having an overcharge protection mechanism, a preferred lithium salt in the solid electrolyte is LiI (5 to 50% W/W), and a cathode of: CoS_2 , FeS_2 , NiS_2 or mixtures of these.

The Li/ FeS_2 system was claimed to be non-rechargeable in non-aqueous solutions. We found that Li/PE batteries with FeS_2 , NiS_2 , CoS_2 cathodes which use $LiAsF_6$ or

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LiBF₄ salts can not be effectively recharged. However when the MX salt (or its mixture) is used they can be effectively charged and recharged and moreover they have an internal electrochemical overcharge protection mechanism. Therefore they can be stacked in series to build a bipolar battery.

In order to use high voltage cathodes such as lithiated manganese dioxide or lithiated vanadium oxides a LiCl based CSE must be used.

According to one embodiment of the invention there can be provided a multi-cell battery having a bipolar stack configuration, of the sequence: electronically conductive inert substrate, anode, CSE separator and cathode. It can consist of several plates or strings of cells in parallel configuration.

Such batteries can have thin (up to 200 micron each) electrodes, substrate and separator layers.

Preferred CSEs essentially comprise in combination a high surface area matrix of small particles, preferably oxides, generally below about 1 micron, and preferably in the 0.05 to 0.2 micron range, which are coated with a thin coating of suitable alkali metal salts or its solid complex with the polymer and either an organic elastomer, which is an electronic insulator, adapted to form a complex with the metal salt coating the matrix particles, or a suitable hydrocarbon polymer, preferably polyalkylene polymer. The novel composite solid electrolytes (CSEs) have the following advantages: (1) Alkali metal transference number greater than 0.5. (2) The salt anions are fully compatible with lithium and other alkali metals. (3) The novel CSE

~~can be used to make a rechargeable Li/FeS₂, Li/NiS₂,~~
Li/CoS₂ secondary cells which have an internal
electrochemical overcharge protection mechanism.

As a result of (2) and because of other factors, the interfacial Li/CSE resistance is very low (up to one order of magnitude lower than that of prior art PEs) and very stable. The CSEs have a better mechanical strength and can be used above 100°C, if needed. The ionic conductivity of the CSEs at 20-100°C is similar to, or better than, that of prior art PEs.

The novel CSEs allow development of alkali metal batteries with a power density above 100 W/Kg and an energy density of above 100 Wh/Kg (based on full charged battery). It is believed that there provided coated very small Al, Mg or Si oxide particles (0.05 μ m) with a thin MX layer or its solid complex with the polymer and which particles are bound together by means of a polymer electrolyte or polymer binder. It is believed that these particles, retain the interfacial and the vacancy conduction mechanism responsible for near unity transference number of M while the film retains the flexibility of the polymer. This results in Composite Solid Electrolytes (CSEs) with a transference number above 0.5, preferably up to close to unity. Another benefit of the oxide matrix is the immobilization of the molten phase when heating the CSE to above the melting point of the polymer. In addition to its binding, the PE provides ionic conductivity between the oxide particles coated by MX (or bridge between them). Others have used Al₂O₃ (2) or alumina based oxides to improve the mechanical strength of polymers but the Li transference number was low as they used a too low salt concentration, the wrong salt and too large Al₂O₃.

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particles. It is believed that the oxide effect on MX is a surface effect and not a bulk one.

PEO is thermodynamically non-compatible with alkali metals. However, it is kinetically stable and reacts only slowly. There exist more stable binders which can be used, such as polypropylene, polyethylene, polystyrene, polyacrylates, polyethylene glycol cross-linked and other hydrocarbon based polymers.

According to this invention the properties of the previously mentioned CSE is improved, and the concept is extended to other CSE films containing a variety of inorganic oxides and a variety of alkali metal salts.

The following alkali metal salts conduct in the solid state alkali metal cations through a metal cation vacancy mechanism, especially at the grain boundaries. The anions are totally compatible with the metal and the MX are electronic insulators: MF, MBr, MCl, MI, M_2S (M is Li, Na, K). One purpose of this invention is to make good composite solid electrolytes which have t_m + close to unity and good ionic conductivity. This may be accomplished by:

1. Fabricating this CSE film of small oxide particles of multivalent elements such as the oxides of Al, Si, B, Ca, Sr, Mg or mixed oxides coated with M-halides. These oxides should be electronic insulators and compatible with the alkali metal. It is believed that the multivalent elements in the oxide induce M^+ vacancies in the MX salt or in the solid MX-Polymer complex, thus increasing the cationic conductivity;

2. Adsorbing multivalent metal compounds on the surface of fine MX particles. These compounds should be compatible with the alkali metal. For example: SrX_2 , CaX_2 , etc.

3. Adsorbing on the fine MX particles one or a few monolayers of organic molecules (compatible with alkali metals) which preferentially complex the alkali metal cation (and not the X anion). Such as di-, tri-, tetra-Glymes, polyethylene glycol dimethyl ester, Crown ethers, PC (propylene carbonate) etc. These M-complexing agents are believed to help by inducing M vacancies in the solid MX by dissolving M from the MX surface and by connecting the solid particles by a liquid electrolyte phase,

4. Optionally doping these salts with multivalent cations such as Al^{3+} , Ca^{2+} , Ba^{2+} , Mg^{2+} , B^{3+} , Sr^{2+} .

Generally the dispersed particles will be of a size below $1\text{ }\mu\text{m}$, and preferably in the 0.01 to $0.2\text{ }\mu\text{m}$ range. The coating of alkali halide (MX) applied to the particle surface is generally of the order of some hundreds of Angstroms (A), up to about $1\text{ }\mu\text{m}$.

In most cases the non-complexing polymer will comprise not more than about 20 weight per cent of the composition, and preferably about 5 to 15 weight per cent.

A preferred component of ion complex forming polymer is PEO (polyethylene oxide). There can also be used polypropylene oxide and the like. Such complex forming polymers are used to up about 80% of the composition.

Mobility of the ionic

species occurs mainly in the amorphous phase. That is why it is preferable to add to PEO electrolytes polymers with low glass transition temperature, such as polyacrylate, polymethacrylate, polybutylmethacrylate, polymethylstyrene and their derivatives.

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Addition of these components increases the rate of solubility of polymers in the casting solution: thus a homogeneous slurry can be readily received in a smaller volume of low dielectric constant organic solvent (for example - acetonitrile).

In addition at certain composition range the mixed polymer films have a better mechanical stability and higher ionic conductivity. For example in Table I CSE which accorded this invention contains PMMA (No. 2) has five times greater ionic conductivity than has the state of the art CSE (No. 1) which does not contain PMMA.

The invention is further illustrated with reference to the enclosed schematical drawings, not according to scale, in which:

Figure 1 is a side view of an electrochemical cell containing an electrolyte of the invention;

Figure 2 illustrates arrangements for measurements of the properties of the novel electrolytes;

Figure 3 illustrates a setup for conductivity measurements;

Figure 4 illustrates a setup for battery tests.

Figure 5 illustrates Bode plot of CSE ($t_+ < 1$)

Figure 5a illustrates Bode plot of CSE ($t_+ = 1$)

Figure 6 illustrates Nyquist plot of CSE ($t_+ < 1$)

Figure 7 illustrates charge-discharge plots of Li/CSE/TiS₂ cell.

Figure 8 illustrates charge-discharge plots of Li/CSE/FeS₂ cell.

Figure 9 illustrates plot of capacity vs number of cycles for Li/CSE/FeS₂ cell.

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Figure 10 illustrates charge-discharge plots of Li/CSE/FeS₂ cell.

Figure 11 illustrates charge-discharge plots of Li/CSE/FeS₂ cell.

As shown in Fig. 1, the experimental battery incorporating a solid electrolyte of the invention comprises a glass container 11, provided with outlet 12, connected to a vacuum pump; 13 is the glass cover of the said vessel, with O-ring 14 between them. In the glass container 11, there are positioned a cell 15 illustrated in Figure 2, connected by wires 16 to tungsten rods 17 which pass through the cover 13.

The unit designated as cell 15 in Figure 1 is illustrated in Figure 2 and it comprises s.s. holder 21, through which there pass screws 22 and 23, there being provided stainless steel (s.s) electrodes 24 and 25, held by ceramic (Macor) holder 26 and 27, pressed together by spring 28. This assembly is adapted to hold in place the devices illustrated in Figures 3 and 4, at position 29. The setup for conductivity measurements of Figure 3 comprises s.s. electrodes 31 and 32, lithium foils 33 and 34, and the conductive polymer 35.

The setup for battery tests illustrated in Figure 4 comprises s.s. electrodes 41 and 42, lithium foil 43, conductive polymer 44 and cathode 45.

The transfer number of the lithium cations in the film can be calculated according to Equation 1 (Weston and Steele - Solid State Ionics, 7, 75 (1982)):

$$t_{Li+} = \frac{R_b}{R_a + R_b}$$

where R_d is the diffusion resistance of the film (PE or CSE) (1) and R_b is the film resistance measured with a AC technique. R_b and R_d can be calculated from Bode Plot (Figure 5) or Nyquist Plot (Figure 6) using a small AC perturbation.

Measurements of the film produced according to Example 1 showed that R_d is practically zero, and thus $t_{L1}^- = 1$.

The invention is illustrated with reference to the following Examples, which are to be understood in a non-limitative manner.

Example 1

To prepare CSE film with the composition LiBr-P(EO)_4 with 6% (v/v) SiO_2 , we used: 304 mg LiBr, 614 mg PEO and 88.3 mg SiO_2 .

The preparation of the composite polymer film was entirely performed in a dry-box, kept under an argon controlled atmosphere and having an average water content less than 20 ppm. The known quantity of the lithium salt was dispersed in AR dry acetonitrile (approximately 50 cc per gram of polymer).

PEO $\text{MW}=5 \times 10^6$ was added to the mixture. After stirring at room temperature for about 4 hours a homogeneous dispersion was obtained and high surface area SiO_2 was added. For quick and for best dispersion of the oxide in the slurry an ultrasonic bath was used. Then partial evaporation was allowed and the viscous (about 20 cm² volume) slurry was

cast on a Teflon tray (64 cm² area) to form a film. The film was dried about 3 hours at room temperature under vacuum and later it was dried at 120°C under vacuum. The final product was a film about 100 μm thickness.

The properties of the polymer materials have been examined in a suitable electrochemical cell (Figure 1).

The electrode area was 1 cm² and the stack pressure was 8 Kg/cm². The polymer conductivity was measured by non blocking electrode (Li), by AC conductivity measurements at the frequency range of 10 MHz to 1 MHz. The conductivity of this film at 130°C is 10⁻⁵ ohm⁻¹cm⁻¹.

Example 2

A CSE was prepared as described in Example 1, starting with the following composition: 436mg LiI; 84.2mg AlI₃; 455 mg PEO and 134mg Al₂O₃ (0.05μ particle size); 76mg triglyme. Its conductivity at 110°C is 1 x 10⁻⁴ ohm⁻¹ cm⁻¹, t_{Li^+} as calculated from equation 1 is about 0.8. In this film the EO to LiI ratio is about 3:1 and the volume percent of Al₂O₃ is 6%.

Example 3

A 100μ thick CSE was prepared as described in Example 1. It contained 50% of PEO of low molecular weight (MW 100,000) and 50% of PEO of high molecular weight (MW 5,000,000).

The casting slurry contained: 336mg LiI; 115 mg PEO (MW 100,000); 115 mg PEO (MW 5,000,000); 134 mg 0.05μ Al₂O₃ and 220 mg EC. The EO to LiI ratio is about 3:1 and the volume percent of Al₂O₃ is 6%.

The conductivity of film at 90°C was $7 \times 10^{-5} \text{ ohm}^{-1} \text{ cm}^{-1}$. Its cationic transference number is close to unity.

Example 4

A CSE was prepared as described in Example 1: a few films with different compositions were produced and measured with the use of AC technique. The results are summarized in Tables 1 and 2. Sample 1 is a state of the art film composition and has a low conductivity. Sample 1 with no PMMA and EC was described by Nagasubramanian et al. Proceeding of ECS Toronto meeting, October, 1992. We added to this composition PMMA and EC which improves film homogeneity and conductivity. Sample 2 has $t_{Li^+} = 0.76$ at 120°C and $t_{Li^+} = 1$ at 90°C.

Example 5

For comparison a CSE without EC was prepared as described in Example 1. The composition of the casting slurry was: 466mg LiI, 458mg PEO, 104mg PMMA and 134mg Al_2O_3 . Its conductivity at 120°C was $9.10^{-5} \text{ ohm}^{-1} \text{ cm}^{-1}$. Lithium cations transference number is 0.6.

Example 6

A CSE was prepared as described in Example 1. A slurry containing PMMA with the following composition $(LiI)_1 P(EO)_3 P(MMA)_{0.3} (EC)_1$ was cast. The casting slurry contained: 336mg LiI, 330mg PEO, 75mg PMMA, 220mg EC and 85mg SiO_2 . The volume concentration of SiO_2 in the final polymer was 6%. Its conductivity at 120°C was $2.10^{-5} \text{ ohm}^{-1} \text{ cm}^{-1}$; Lithium cations transference number is close to unity.

Example 7

A 200 micron thick CSE with a composition of 6% (v/v) MgO (LiI)₁ P(EO)₃ P(MMA)_{0.25} ethylene carbonate (EC)₁ was produced following the procedure of Example 1. The conductivity of it at 120°C was $2 \cdot 10^{-4} \text{ ohm}^{-1} \text{ cm}^{-1}$. The value of lithium cations transference number was over 0.5.

Example 8

A composite solid electrolyte film, containing poly(butylacrylate) with low glass transition temperature of -49°C was obtained according to procedure of Example 1. The conductivity of a film with the composition of 6% Al₂O₃LiI(PEO)₁₀ P(BA)_{0.5} EC₁ at 120°C was $9 \times 10^{-4} \text{ ohm}^{-1} \text{ cm}^{-1}$. A 100 μm thick CSE contained: 136mg LiI, 452mg PEO, 65mg PBA 90mgEC.

Example 9

A CSE was prepared as described in Example 1 A 100 micron thick film with the composition: 93 mg LiI, 183 mg PEO, 42 mg PMMA, 145 mg polyethylene glycol dimethyl ether (PEGDME) was produced. Its conductivity at 120°C is $1 \times 10^{-3} \text{ ohm}^{-1} \text{ cm}^{-1}$ and its t_{Li^+} is over 0.5.

Example 10

A 100 μm thick film doped with CaI₂ was obtained following the procedure of Example 1. Its composition was 460mg LiI, 30mg CaI₂, 450mg PEO, 120mg PMMA, 200mg EC and 134mg Al₂O₃ (6% v/v). Its conductivity at 130°C was $10^{-4} \text{ ohm}^{-1} \text{ cm}^{-1}$.

Example 11

A battery was assembled. It consisted of: a lithium anode, 150 μ thick CSE separator of composition of Example 4, and a 100 μ thick cathode with the following composition: 50% by volume CSE and the rest TiS_2 particles (fine particles). The cathode foil was prepared by dispersing TiS_2 particles (less than 50 μ in size) in $\text{P(EO)}_{2.5}$ - LiI - $(\text{PMMA})_{0.25}$ - Al_2O_3 - acetonitrile solution and casting it into a Teflon mold. Later the cathode foil was dried at 120°C under vacuum. This battery was discharged at 120°C at $I = 0.10 \text{ mA cm}^{-2}$ for 20 hours and then charged at 0.06 mA cm^{-2} for 20 hours, (Figure 7) average discharge voltage was 2.1 V and average charge voltage was 2.3 V.

Suitable alloys anodes which give similar results are alloys of the Li-Al-Mg type and of the Li-Al type. Good results were obtained with an alloy of (percent by weight) of 2 to 30% Li, at least 20% Al and at least 5% Mg. Good results were also attained with a LiMg alloy containing at least 20% magnesium.

Example 12

A battery was assembled as described in Example 11 except that the CSE composition was LiI P(EO)_3 $\text{PMMA}_{0.3}$ 6% v/v MgO . It was cycled at 120°C in the following regime: discharged at 0.05 mA/cm^2 for 10 hours and charged at 0.05 mA/cm^2 for 10 hours. It had a similar average discharge and charge voltage as in Example 11.

Example 13

A Li/CSE/FeS₂ cell was assembled. It consists of: a lithium anode, 150μm thick CSE separator of composition of Sample 2, Table 1, Example 4 and a 50μ thick composite FeS₂ cathode which consists of 60% by volume CSE and the rest FeS₂ fine powder. The cathode was prepared as described in Example 11. The current collector for the cathode was a 2mm thick graphite disc. $i_a=0.1 \text{ mAcm}^{-2}$, $i_c=0.05 \text{ mAcm}^{-2}$. The cell was held under a spring pressure of 12kg/cm² and hermetically sealed in a glass vessel. It was cycled at 120°C (Fig. 8). Charge and discharge capacities were similar and stable for over 20 cycles. This cell could be overcharged 100% with almost no loss in capacity.

A similar Li/CSE/FeS₂ cell was assembled. It consists of a lithium anode, 100μm thick CSE separator of composition of Sample 4, Table 1, Example 4 and 50μm thick composite FeS₂ cathode which consists of 65% by volume the same CSE and the rest was fine FeS₂ powders.

The cell was held under a spring pressure of 5kg/cm². It was cycled at 135°C at 0.05 mAcm⁻² for charge and 0.3 mAcm⁻² for discharge.

Cut off voltages were 1.1V for discharge and 2.15V for charge. The cell delivered above 2mAh per cycle for over 30 cycles (Fig. 9). It could be overcharged at 2.4V for more than 30% with no loss in capacity (Fig. 10). Similar results were obtained when the cell was assembled with a cathode in the discharged state i.e. by a cast of a mixture of fine Fe and Li₂S powders instead of FeS₂ powder.

Example 14

A battery was assembled as described in Example 13. The cathode was 50 μm thick composed of FeS_2 35% (v/v) and 65% (v/v) CSE, CSE thickness was 100 μm . The CSE casting slurry contained 122 mg $\text{Li}(\text{CF}_3\text{SO}_2)_2\text{N}$, 268 mg Al_2O_3 , 561 mg PEO 40 mg, PMMA and 71mg EC. Stack pressure was 4kg/cm². The cell was cycled at 130°C between 1.1 to 2.45 V for over 10 cycles at 0.3 mA discharge current and 0.05 mA charge current. It delivered 2.3 mAh/cm².

Example 15

A battery was assembled as described in Example 13 with stack pressure of 3kg/cm². The CSE composition was $(\text{Li}(\text{CF}_3\text{SO}_2)_2\text{N})_{0.7} (\text{LiI})_{0.3} \text{P}(\text{EO})_{16}, \text{P}(\text{MMA})_{0.5}$ and 9% by volume Al_2O_3 . It was cycled at 130°C at discharge current at 0.3mAcm⁻² and charge current of 0.045mA/cm⁻². The cell delivered over 2.5mAh for over 10 cycles (Fig. 11).

Example 16

A battery was assembled as described in Example 15 with 100 μm thick CSE separator of Sample 4, Table 1, Example 4. The anode was 1 mm thick sodium metal. This cell was cycled at 80°C between 0.8 and 2.4V at discharge and charge currents of 0.03 mA. The cell delivered about 0.1mAh for several cycles.

Example 17

A battery was assembled as described in Example 13. It consists of a NiS_2 cathode instead of FeS_2 cathode, 100 μm thick CSE of composition of Sample 4, Table 1, Example 4. The stack pressure was 4kg/cm². It was cycled for over 10 cycles at 135°C between 1.1 and 2.15V at 0.1 mA discharge and 0.05 mA charge. It delivered 1.5 mAh/cm².

Example 18

A battery was assembled and tested as described in Example 17, except the cathode is $\text{CoS}_{2.5}$. It delivered about 1.8mAh for over 5 cycles.

Example 19

A Li/CSE/LiMnO₂ battery was assembled. It consists of Li anode, 150 μm thick CSE and 100 μm thick composite LiMnO₂ cathode. CSE film was casted from a slurry consisting of 78mg LiCl; 483mg PEO; 46mg PMMA; 134mg Al₂O₃; 161mg EC. The cathode consist of 65% v/v CSE, and 25% v/v LiMnO₂ and 10% v/v carbon powder and prepared as described in Example 11. The cell was cycled at 135°C between 2.5 to 3.5V at current of 0.02mA for both charge and discharge. It delivered 0.1 mAh for over 5 cycles.

Example 20

A CSE was prepared as described in Example 1: 100 μ thick film with the composition: 113mg LiI, 60mg KI, 480 mg PEO, 60mg PMMA, 107mg EC, 301mg Al₂O₃, was produced. Its conductivity at 120°C is $1.5 \times 10^{-3} \text{ Ohm}^{-1} \text{ Cm}^{-1}$.

Example 21

A CSE with the composition 25% (v/v) PAN, 25% v/v EC; 25% v/v PC, 24% v/v Dioxolane 6% Al_2O_3 (v/v) and 1.0 M LiCF_3SO_3 . The conductivity of this CSE at room temperature was $1.1 \times 10^{-3} \text{ Ohm}^{-1} \text{ cm}^{-1}$.

Example 22

A cell was assembled with carbon anode consisted of 35% (v/v) fine powder of petroleum coke and 65% (v/v) CSE with the following composition:
0.3M $\text{Li}(\text{CF}_3\text{-SO}_2)_2\text{N}$, 0.7M LiI , 30% v/v, PEO 30% v/v PC, 30% v/v DEC, 6% (v/v) Al_2O_3 . The counter electrode was a lithium foil, and the CSE separator between them consisted of the same composition. This cell was charged-discharged at 80°C between the anode composition Li_0C_6 and $\text{Li}_{0.4}\text{C}_6$ for over 50 cycles.

Table 1: The Effect of Composition on Z_T , R_{SEI} , R_{CSE} , σ , t_+ at 120°C

No CSE Composition	Z_T ΩCM^2	R_{CSE} ΩCM^2	Z_D 2CM^2	R_{SEI} ΩCM^2	σ $\Omega^{-1}\text{CM}^{-1}$ ($\times 10^4$)	t_+
LiI-P(EO) ₃ * ⁺	216	168	34	14	1.12	0.83
LiI-P(EO) _{2.5} X _{0.25} * [°]	75	34	11	30	5.9	0.7
LiI-P(EO) ₆ X ₁ [°]	63	20	32	11	5.0	<0.38
LiI-P(EO) ₉ X _{0.5} [°]	57	10	24	23	10.0	<0.3

Sample 1 is state of the art example and not a part of this invention.

* - Thickness 200 μ

° - Thickness 100 μ

+ - 12 Vol % Al₂O₃, the rest have 6%

x - PMMA; X_{0.5} means mole ratio of 0.5 PMMA to LiI.

Samples 2, 3 and 4 contain 1 mole of EC to one mole of LiI in the casting slurry.

Z_D - diffusion impedance, σ -CSE conductivity

Z_T - Total impedance; R_{SEI} : the resistance of the solid electrolyte interphase on the surface of the lithium i.e. the Li/CSE interfacial resistance not to confuse with CSE resistance (R_{CSE})

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Table 2: The Effect of Composition on Z_T , R_{SEI} , R_{CSE} , σ , t_+ at 90°C

No CSE Composition	Z_T ΩCM^2	R_{CSE} ΩCM^2	Z_D 2CM^2	R_{SEI} ΩCM^2	σ $\Omega^{-1}\text{CM}^{-1}$ ($\times 10^4$)	t_+
LiI-P(EO) ₃ * ⁺	806	613	0	194	0.3-0.4	1
LiI-P(EO) _{2.5} X _{0.25} *264		58	0	206	3.4	1
LiI-P(EO) ₆ X ₃ ^o	180	34	>75	70	2.9	<0.32
LiI-P(EO) ₉ X _{0.5} ^o	182	15-20	>68	92	4.5	<0.25

* - Thickness 200 μ o - Thickness 100 μ + - 12 Vol % Al₂O₃, the rest have 6%

Example 1 - not part of this application.

CLAIMS:

1. A composite solid electrolyte (CSE), for use in electrochemical cells, which CSE has an alkali metal transference number larger than about 0.5, which CSE comprises up to about 30 volume-% electronically non-conductive oxide particles, of an average size of less than about 5μ , which are compatible with alkali metals, up to about 70 volume-% of an alkali metal salt MX, where M is an alkali metal and X an anion, which forms a thin coating of said oxide particles or is in contact with these, which salt MX is not reducible by the alkali metal in such cell, up to about 70 volume-% of an electronically insulting elastomer, or hydrocarbon polymer, which elastomer forms a complex with the alkali metal salt, and up to 20 volume-% of an aprotic organic solvent or low MW polymer, excluding the system of polyethylene oxide (PEO), LiI and Al_2O_3 only.

2. A CSE according to claim 1, where the size of the oxide particles is in the 0.05 to 0.5μ size range, a CSE, where the quantity of oxide particles is from 1 to 20 volume-%, the quantity of the alkali metal salt is from 10 to 50 volume-%, and which contains from about 30 to 70 volume-% elastomer and/or hydrocarbon polymer.

3. A CSE according to claims 1 or 2, where the oxide particles are selected from oxides of Al, Si, Mg, Ca, Ba, Sr, and mixtures of any of these which may contain some Li, Na, K, where the alkali metal of the MX salt is selected from Li, Na and K and the anion X is selected from I^- , Br^- , Cl^- , S^{2-} , N_3^- , $CF_3SO_3^-$, $(CF_3SO_2)_2N^-$, and a mixture of any of these, where the elastomer is a polyalkylene oxide, possibly in

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combination with one or more of PMMA, PMA, PBA, PBMA, PAN, a polyalkylene glycol dialkyl ester or ether, where the ratio of the polyalkylene oxide to the other polymers varies from 50: 1 up to 1:1, and an electrolyte in which the polymer is a mixture of polyethylene oxides $P(EO)_n$ where n is a large integer, such that the M.W. is from about 10^5 to 10^7 , and where one MX group is present per 2 to 30 EO groups.

4. A composite electrolyte according to any of claims 1 to 3 containing up to 20 weight-% diglyme, triglyme or tetraglyme or of a crown ether or polyalkylene glycol dialkyl ether (or ester) of MW 500 to 50,000 or an aprotic organic solvent, an electrolyte, where the solvent is propylene carbonate, ethylene carbonate, γ -butyrolactone, DMSO Diethylene carbonate, THF, 2MTHF, Dioxolane, DME or a mixture of solvents.

5. Composite Electrolyte according to any of claims 1 to 4, where the MX salt is doped by up to 10% atomic ratio of a polyvalent cation selected from Al, Ca, Mg, Sr and Ba.

6. A primary or secondary electrochemical battery comprising a solid composite electrolyte according to any of claims 1 to 5.

7. A battery according to claim 6 comprising an anode made from an alkali metal (M) or alkali metal alloy or a M_xC anode, a composite solid electrolyte separator as claimed in any of claims 1 to 5 compatible with the anode, and the cathode which is easily charged and discharged against the relevant alkali metal and current collectors for the anode and the cathode, with the preferred M being Li, where the cathode material is selected from: TiS_2 ,

Li_xMnO_2 ($0.5 < x < 1$), V_2O_5 , CoS_x , NiS_x ($0.8 < x < 4.5$), V_3O_{16} , FeS and FeS_2 , which may be mixed with carbon and CSE composition to form composite cathode.

8. A battery according to claim 6, where the lithium salt in the solid electrolyte is LiI (15-70%) W/W, with up to 50 mol % LiBr , LiCF_3SO_3 , $\text{Li}(\text{CF}_3\text{SO}_2)_2\text{N}$ or their mixture, and the cathode is selected from CoS_x , NiS_x , ($0.8 < x < 4.5$) and FeS_2 .

9. A battery according to claim 6 where the alkali metal is sodium, the CSE contains sodium halide and the cathode active material is selected from FeS , CoS_x ($0.8 < x < 4.5$) and FeS_2 , where the oxide matrix is selected from $\text{Na-Al}_2\text{O}_3$, $\text{Na}^+\text{BAl}_2\text{O}_3$, Al_2O_3 , SiO_2 or MgO , where the salt in the CSE is selected from NaI with up to 50 mol % NaBr , NaCF_3SO_3 , $\text{Na}(\text{CF}_3\text{SO}_2)_2\text{N}$ or their mixture, or the corresponding Li-salts.

10. A multi-cell battery according to claims 6 to 9 having a bipolar stack configuration, of the sequence: electronically conductive inert substrate, anode, CSE separator and cathode having up to 300 micron thick electrodes, substrate and separator layers.

11. A battery according to claims 6 to 10 having several plates or strings of cells in parallel configuration.

12. A secondary battery consisting of: an alkali metal or alkali metal alloy or M_xC ($x < 0.2$) anode; an anode and cathode compatible polymer electrolyte consisting of an alkali metal salt or salts in which the cathode can be charged, an elastomer which can complex this salt; and a cathode or composite cathode which is selected from FeS_2

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or FeS , NiS_x , and CoS_x ($0.8 < x < 4.5$), where the polymer electrolyte and cathode optionally contain salts and oxides defined in claims 2 and 3, and short polymers and solvents defined in claims 3 and 4, with the preferred M being Li and the preferred cathode being FeS_2 .

13. A battery according to claims 6 to 12 where the cell is assembled with a cathode in its discharge state instead of charge state.

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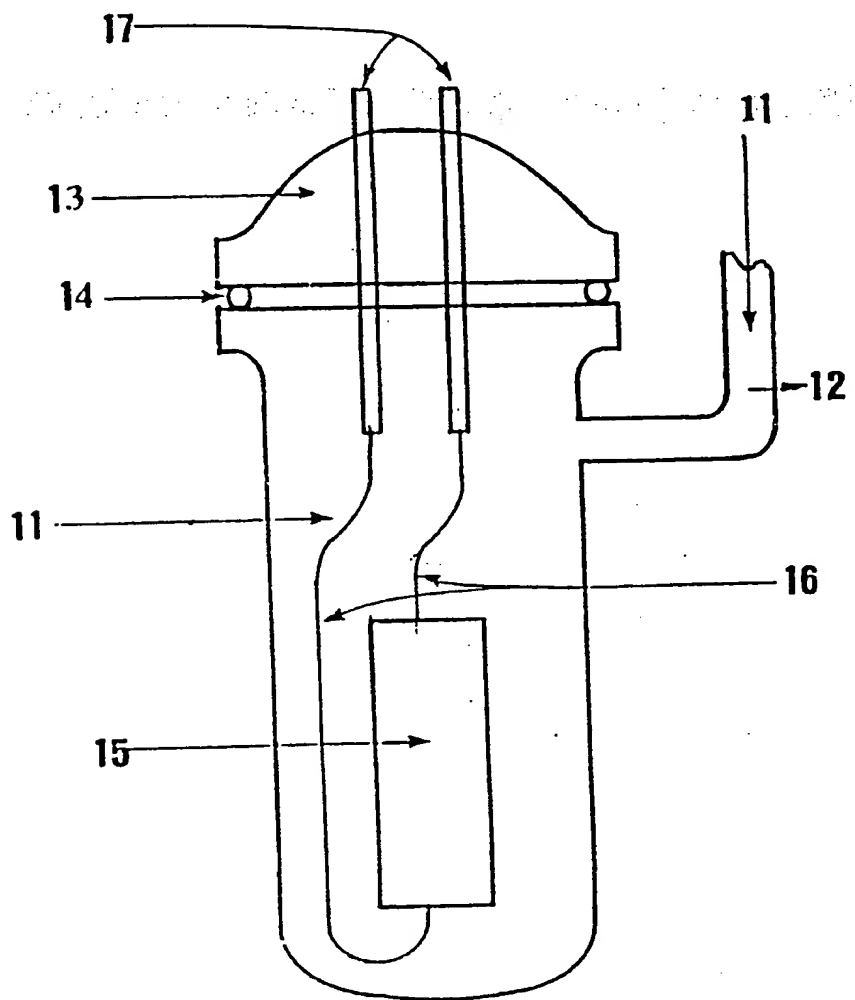


Fig. 1

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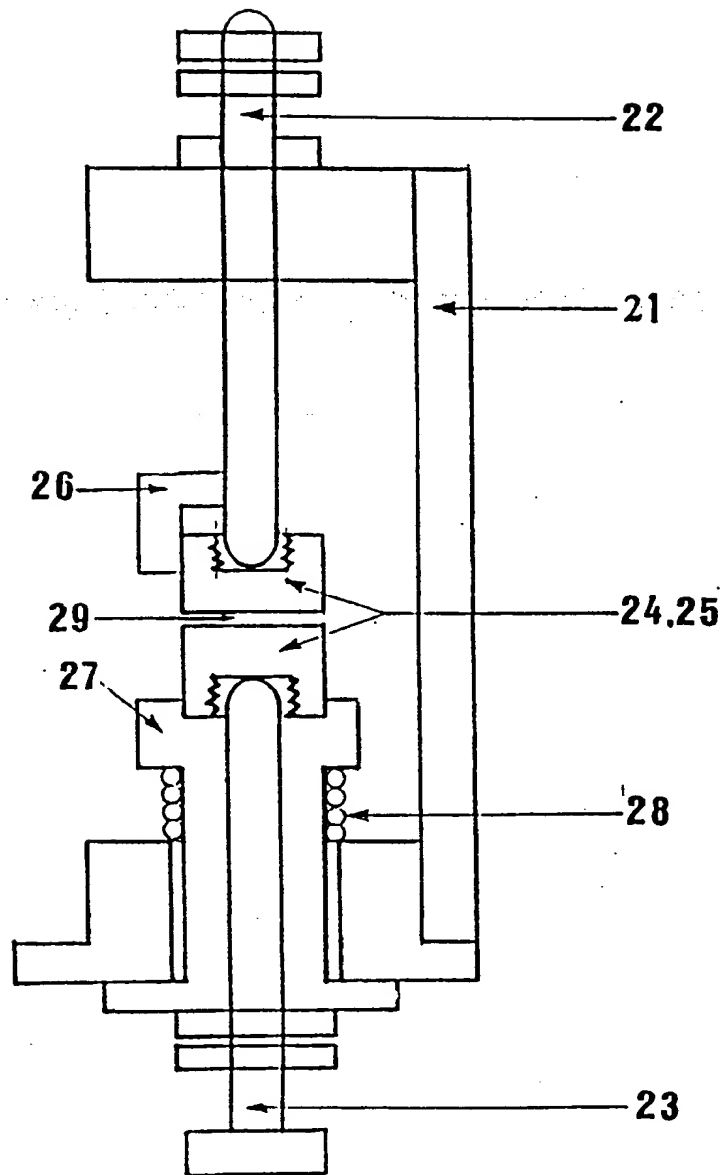


Fig. 2

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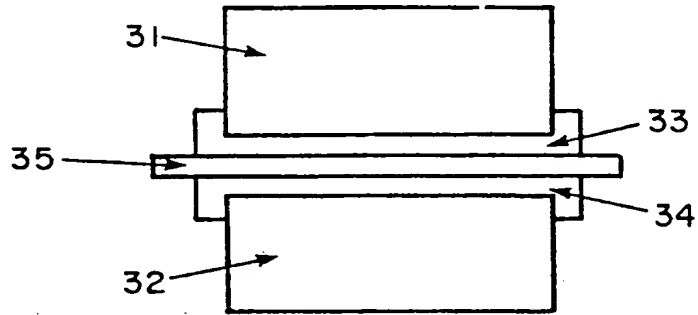


Fig. 3

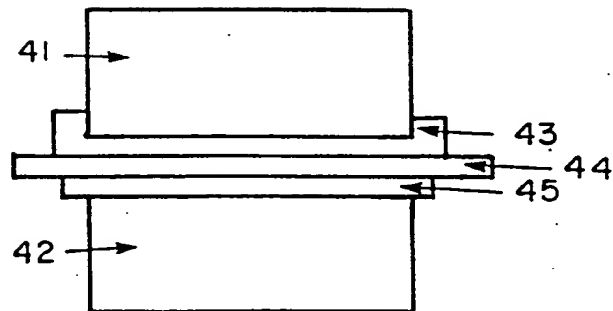
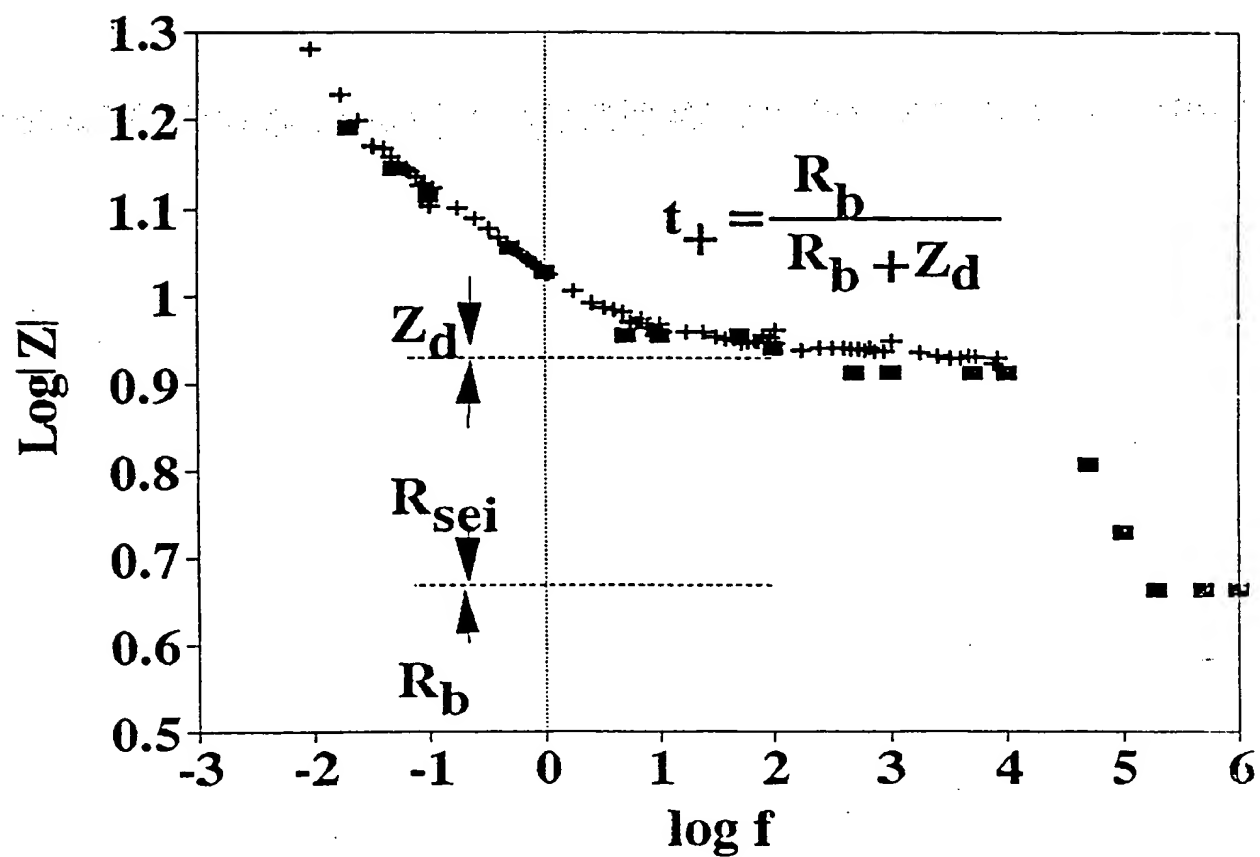


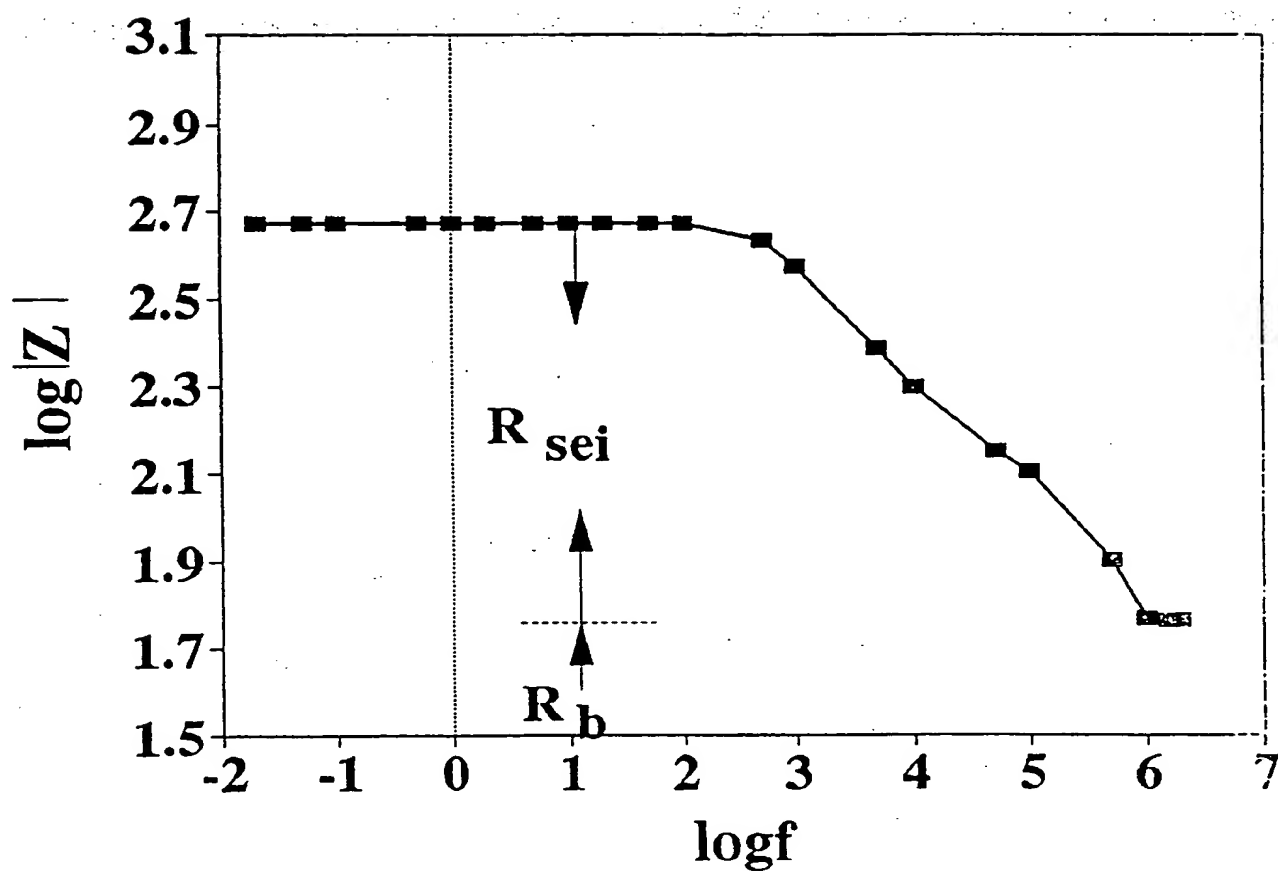
Fig. 4

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Fig.5. BODE plot of CSE ($t_+ < 1$)

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Fig5aBODE plot of CSE ($t_+ = 1$)

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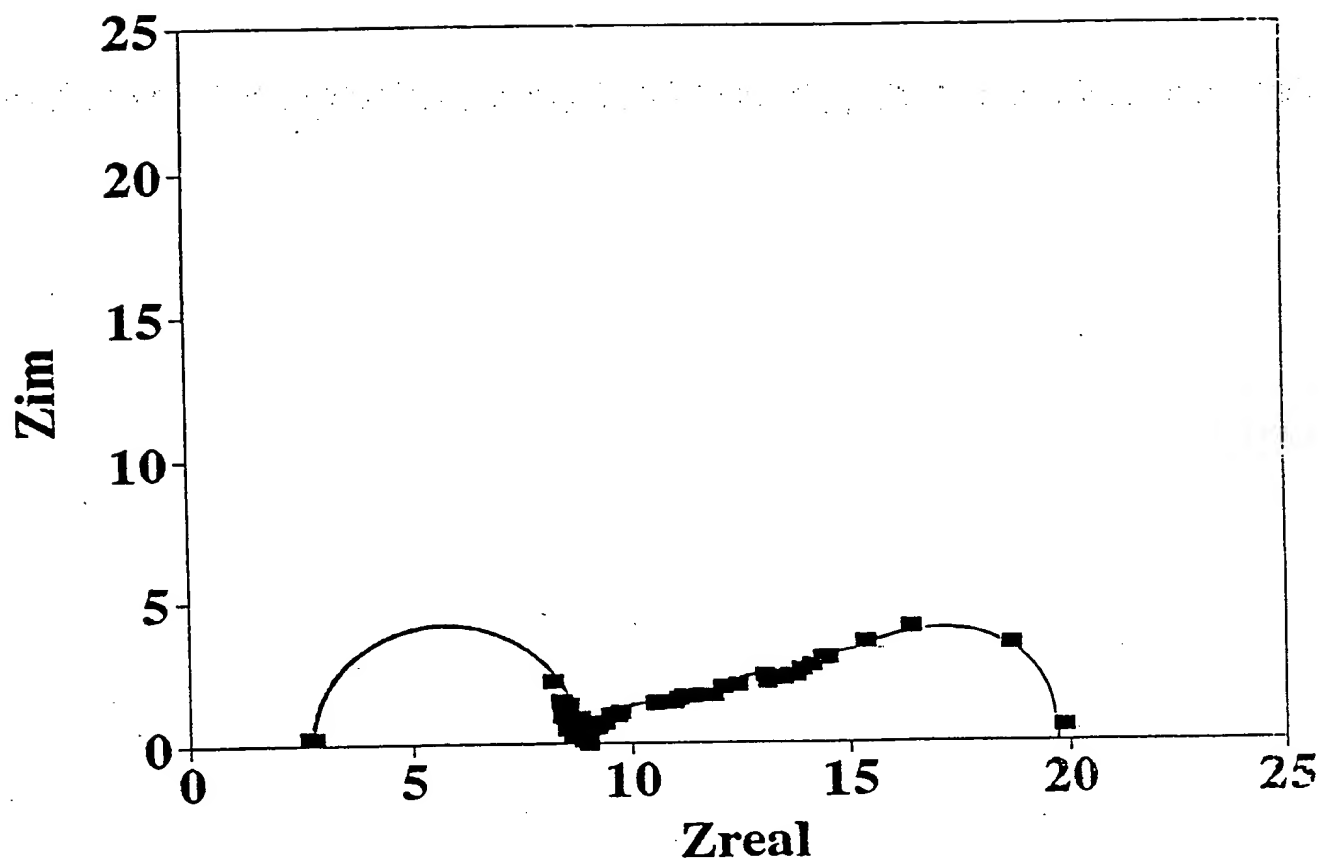


Fig.6 Nyquest plot of CSE ($t_+ < 1$)

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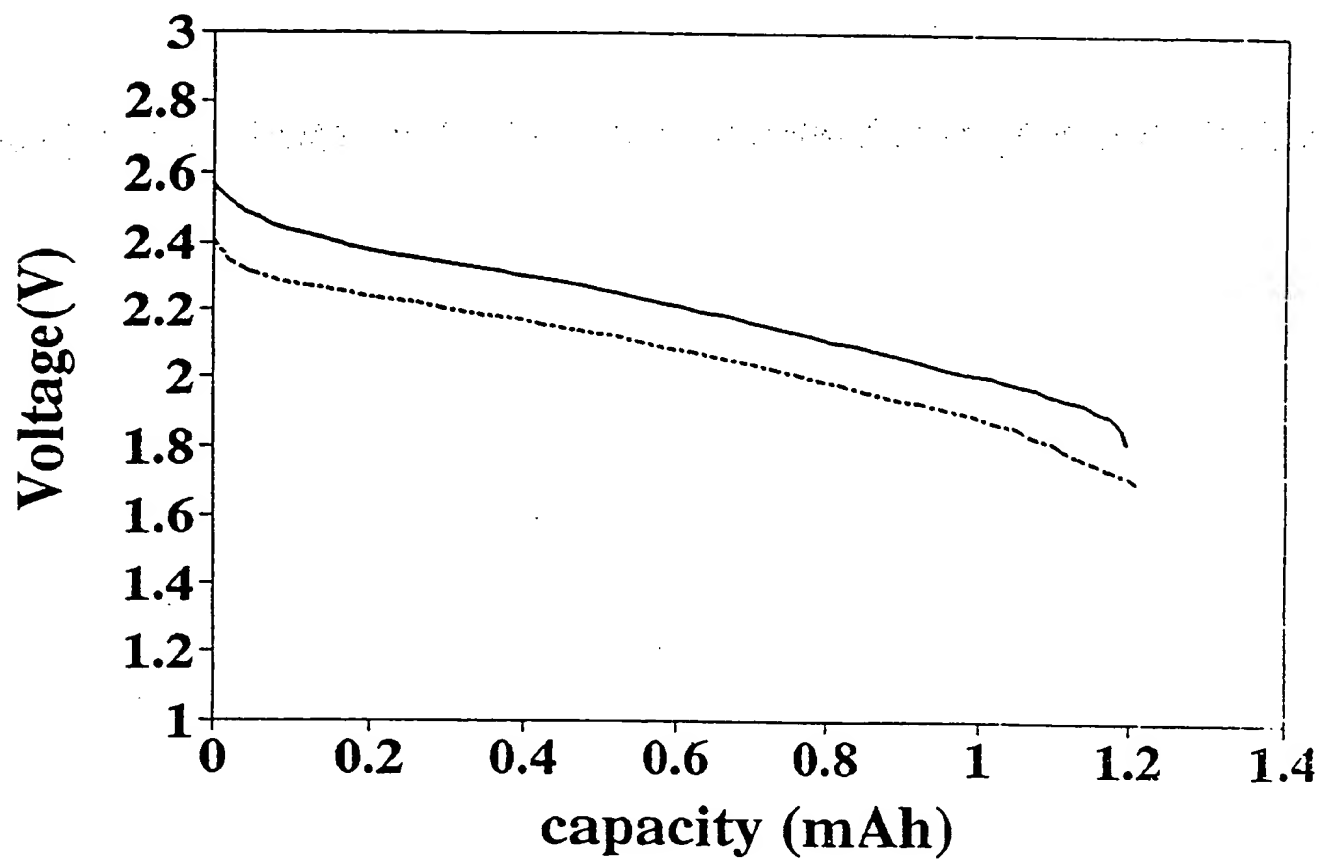


Fig.7. Charge-discharge plots of Li/CSE/TiS₂ cell
at 1.7-2.6 V. $i_c=60\mu\text{A}$, $i_d=100\mu\text{A}$. $T=120^\circ\text{C}$

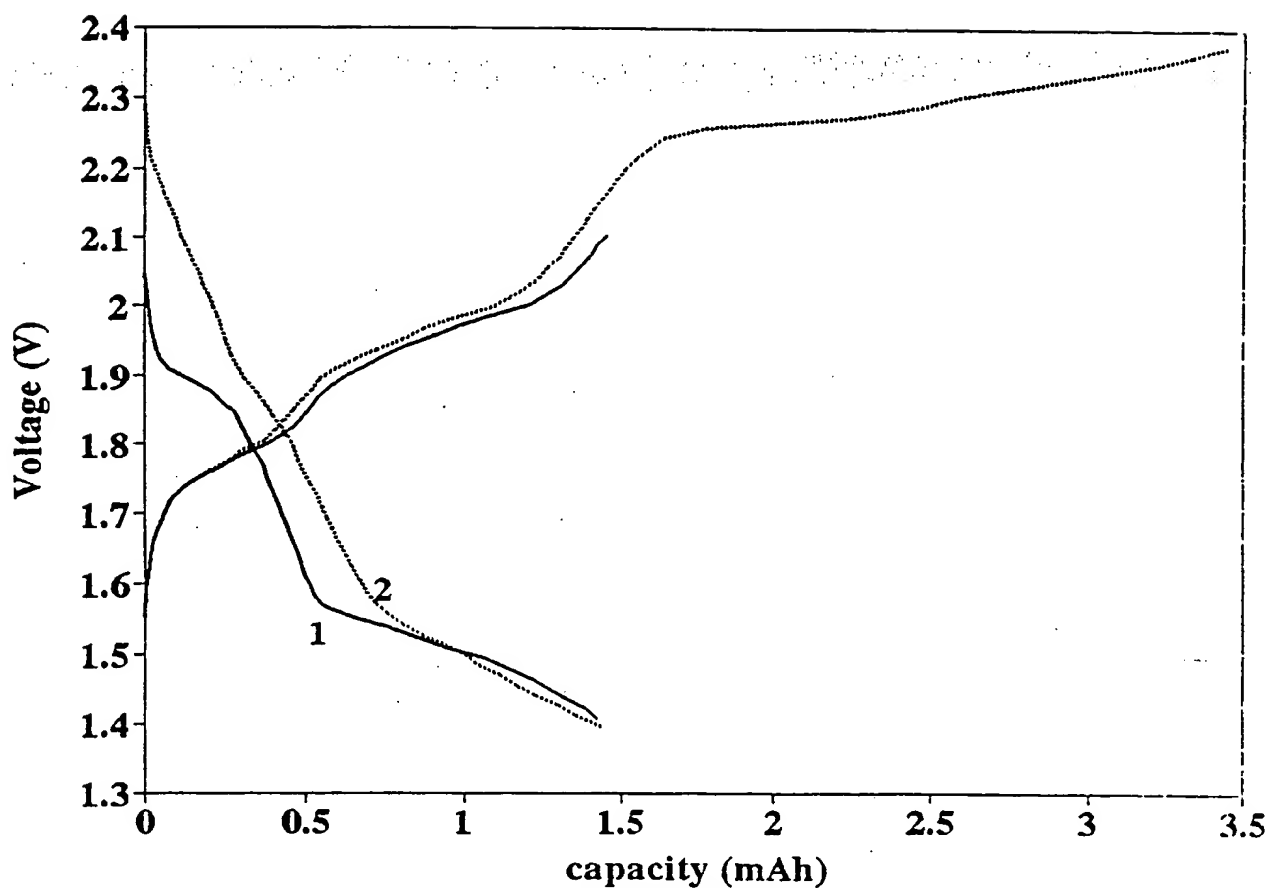


Fig.8 Charge-discharge plots of Li/CSE/FeS₂ cell at 1.1-2.1 V(1), 1.1-2.4 V(2)

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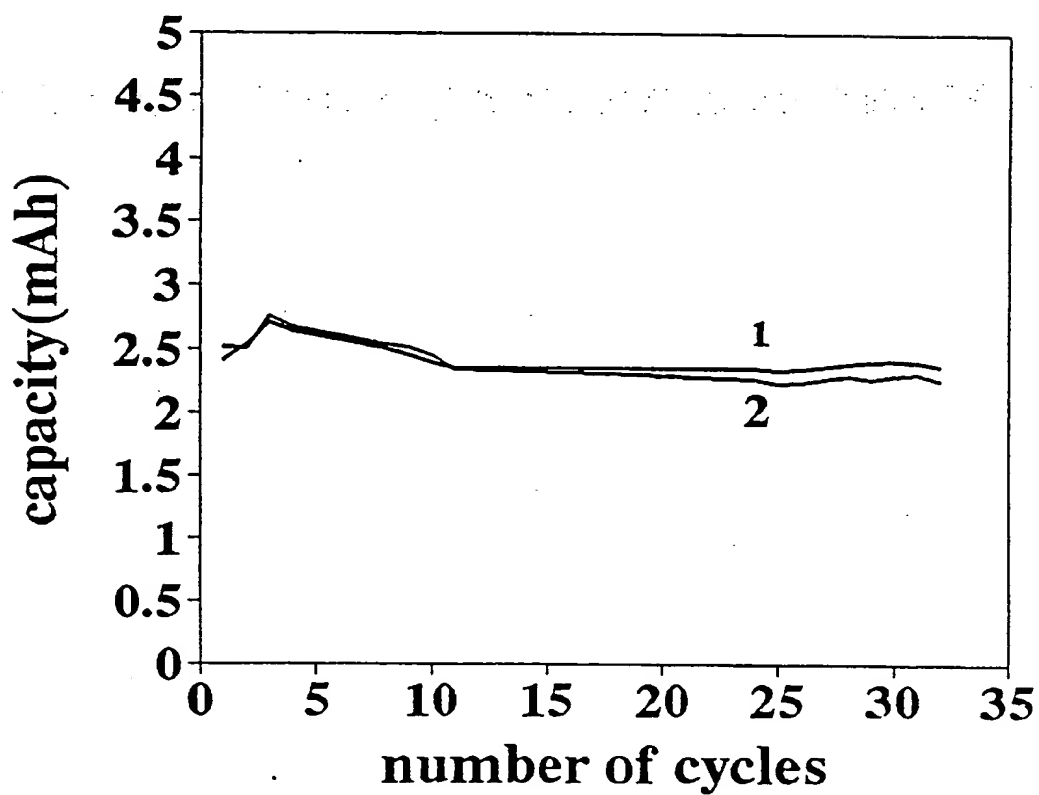


Fig.9 Plot of capacity vs. number of cycles for Li/CSE/FeS₂ cell
1.-charge, 2.-discharge

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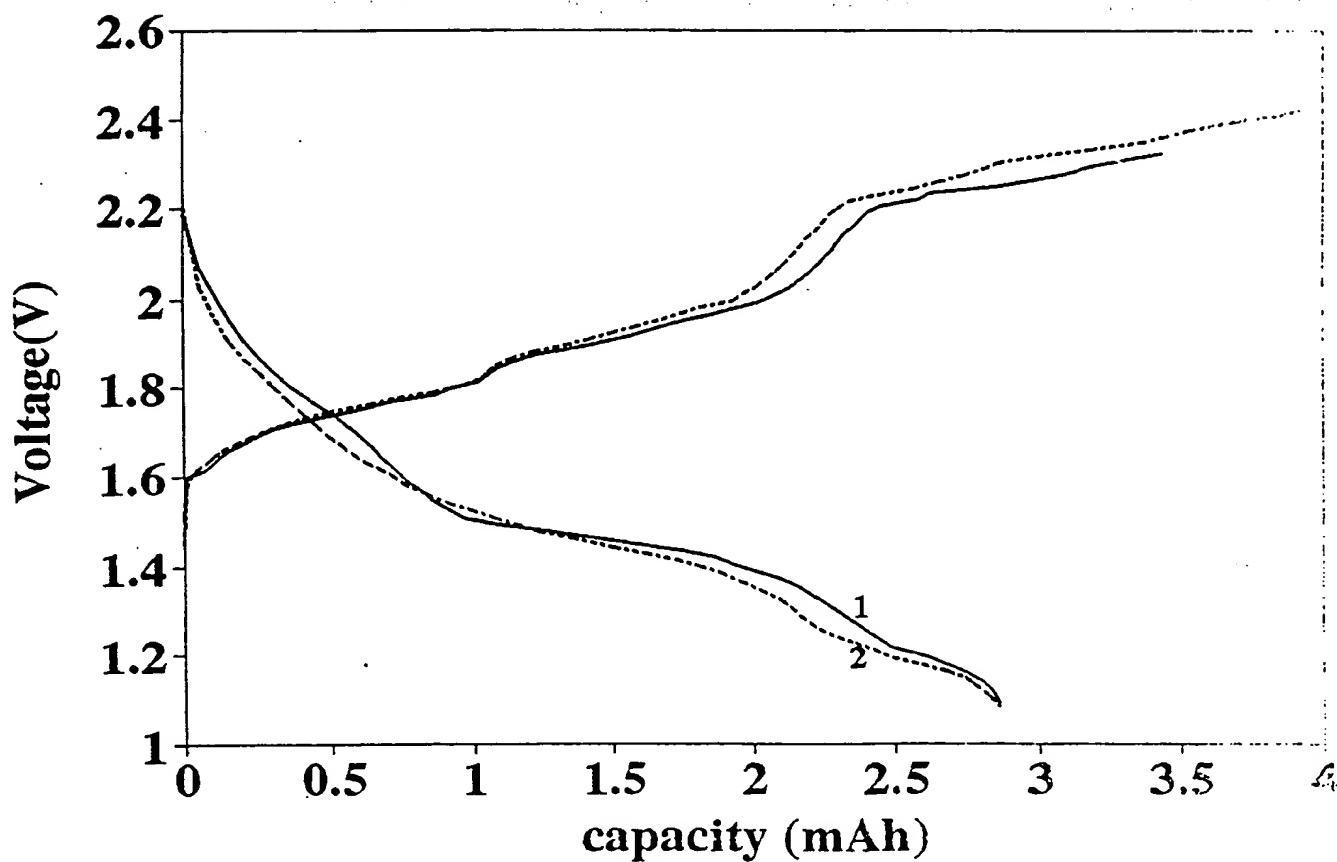


Fig.10. Charge-discharge plots of Li/CSE/FeS₂ cell at 1.1-2.25 V(1), 1.1-2.4V(2)

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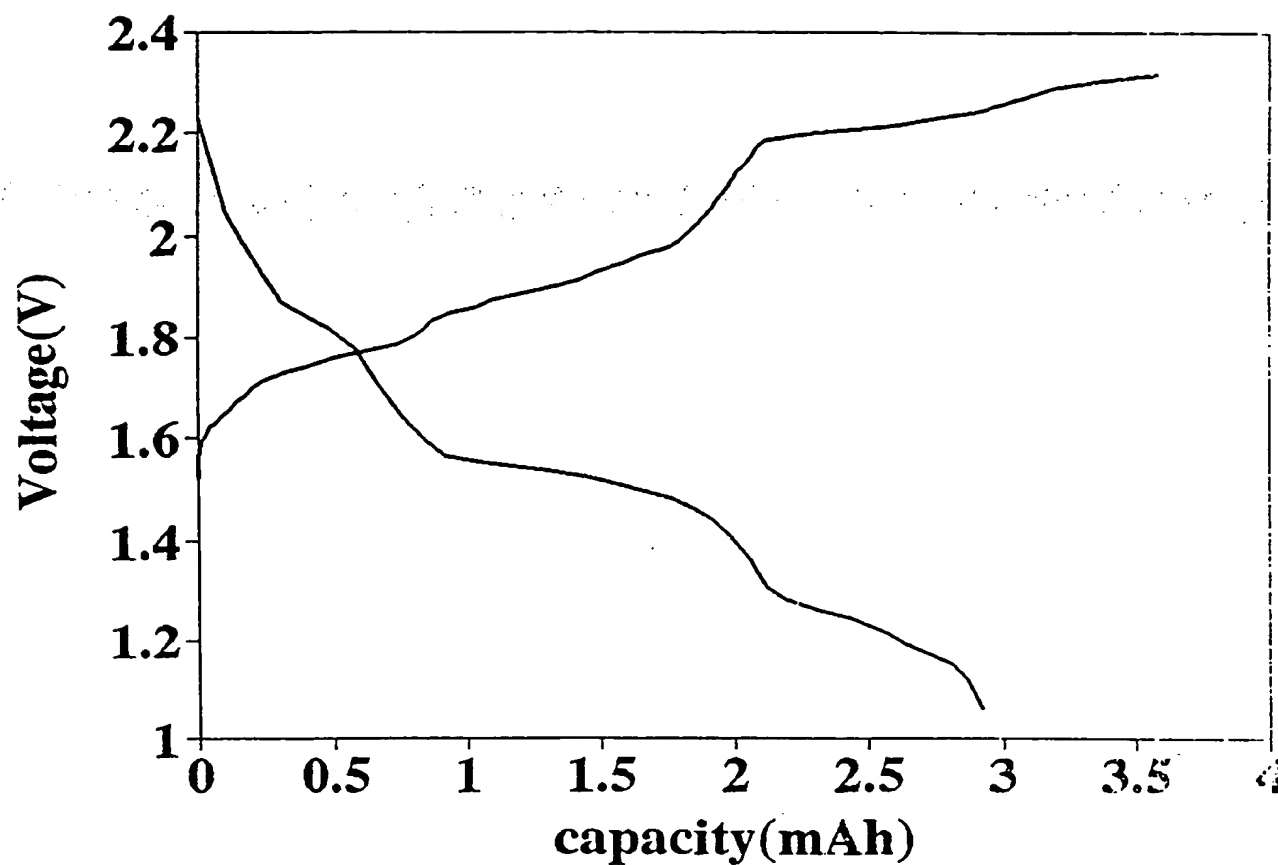


Fig1 Charge-discharge plots of Li/CSE/FeS₂ cell
at 1.1-2.3 V, $i_c=45\mu\text{A}$, $i_d=300\mu\text{A}$, $T=130\text{ C}$

INTERNATIONAL SEARCH REPORT

International application No.

PCT/US94/03800

A. CLASSIFICATION OF SUBJECT MATTER

IPC(5) :H01M 6/18

US CL :429/190, 191, 192

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

U.S. : 429/190, 191, 192

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
Y	US, A, 5,154,987 (HASH ET AL.) 13 OCTOBER 1992, COL. 3, LINES 20-68.	1-13
Y	US, A, 5,030,523 (NEAT ET AL.) 09 JULY 1991, COL. 1, LINES 6-68.	1-13
Y	US, A, 5,102,752 (HOPE ET AL.) 07 APRIL 1992, COL. 2, LINES 41-62.	1-13
Y, P	US, A, 5,204,196 (YOKOMICHI ET AL.) 20 APRIL 1993, COL. 2, LINES 1-68.	1-13
Y	GB, A, 2,157,066 (GLASSE ET AL.), 16, OCTOBER 1985, ALL.	1-13



Further documents are listed in the continuation of Box C.



See patent family annex.

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P document published prior to the international filing date but later than the priority date claimed	

Date of the actual completion of the international search

20 JUNE 1994

Date of mailing of the international search report

JUL 21 1994

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